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The first part of the experiment
 was to determine the rate of
 reaction between hydrogen peroxide
 and iodide ions.

The reaction is as follows:
 $2H_2O_2 + 2I^- \rightarrow 2H_2O + O_2 + 2I^-$
 The rate of reaction was measured
 by the volume of oxygen gas
 evolved over a period of 10 minutes.

The results of the experiment are
 shown in the table below.

Time (min)	Volume of O_2 (cm ³)	Rate of Reaction (cm ³ min ⁻¹)
0	0	0
1	10	10
2	20	20
3	30	30
4	40	40
5	50	50
6	60	60
7	70	70
8	80	80
9	90	90
10	100	100

The graph below shows the
 volume of oxygen gas evolved
 over time.